Exam Key Term Review Matching #1

Matching

a. precision  
g. physical change  
b. accuracy  
h. homogeneous mixture  
c. SI system  
i. heterogeneous mixture  
d. English system  
j. element  
e. Metric System  
k. compound  
f. significant figures  
l. chemical change  

1. any change to matter that causes a new substance to form when the chemical bonds are broken and atoms rearrange
2. an indicator of how correct a measurement in reference to an accepted value
3. the customary system for measurements used in the US for everyday measurements like inches, miles, cups, gallons, ounces, pounds
4. the International System of measurement adopted to standardize units of measure across nations
5. an indicator of how close a series of measurements are to each other
6. measurement system that was developed to use base units and prefixes that change the value of the base by factors of 10
7. all of the digits known for certain by a measurement made plus one estimated digit (the last digit)
8. the simplest form of matter, “building blocks of matter”
13. $G$
14. $K$

may be bonded with ionic or covalent bonds
Exam Key Term Review Matching #2

Matching

a. quantum  j. Neils Bohr
b. ground state  k. cathode ray tube
c. valence electrons  l. gold foil test
d. excited state  m. isotopes
e. octet rule  n. cations
f. group  o. mass number
g. J.J. Thomson  p. anions
h. John Dalton  q. atomic number
i. Ernest Rutherford  r. atomic mass

1. the normal energy level position for electrons
2. a vertical arrangement of elements that share similar properties due to the same amount of valence electrons
3. atoms of the same element with a different number of neutrons
4. atoms that carry a negative charge by gaining electrons
5. generally states that the stability of an atom may be caused by having 8 valence electrons in its’ valence level
6. the minimum amount of energy needed to cause an electron to jump from its ground state to the excited state
7. experiments with this device led to discovery of the electron
8. the number of protons in an atom
9. atoms that carry a positive charge by losing electrons
10. a weighted average mass of the naturally occurring isotopes of an element
11. his experiments with the cathode ray tube led to the development that electrons are smaller particles found in all atoms
12. his experiments with the gold foil test shaped the idea that atoms have a tiny nucleus that is dense and carries a positive charge
13. describes the condition for an electron that has absorbed a quantum and jumped to a higher energy position
14. English school teacher who developed mathematical math ratios of mass as the first evidence to support theory about the existence of atoms
15. the total number of protons and neutrons in an atom
16. his ideas regarding relationships between light energy released and energy level position of electrons led to the development of the Planetary model of atom showing electrons positioned in discreet energy levels
17. the electrons located on the highest occupied energy level
18. experiment that led to the discovery of the tiny dense nucleus of an atom